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PROCESS FOR PREPARING SHORT CHAIN**ALKYL AROMATIC COMPOUNDS**10 **BACKGROUND OF THE INVENTION**

This invention relates to a process for preparing short chain alkyl aromatic compounds by alkylating an aromatic compound with a relatively short chain alkylating agent, employing a binder-free molecular sieve, as the alkylation catalyst.

15 Alkylation is one of the most important and useful reactions of hydrocarbons. Lewis and Bronsted acids, including a variety of natural and synthetic zeolites, have been used as catalysts. Alkylation of aromatic hydrocarbon compounds employing certain crystalline zeolite catalysts is known in the art. For example, U.S. Patent No. 4,992,606, which is incorporated herein by reference, describes the use of an alumina-
20 bound zeolite catalyst, namely an alumina-bound MCM-22 catalyst, in preparing short chain alkyl aromatic compounds, such as cumene.

SUMMARY OF THE INVENTION

It is an object of this invention to provide an improved process for the alkylation
25 of an aromatic compound with a relatively short chain alkylating agent to produce a short chain alkyl aromatic compound, employing an alkylation catalyst comprising a binder-free molecular sieve.

It is a particular object of the present invention to provide a process for alkylating
30 benzene with propylene to produce cumene, or with ethylene to produce ethylbenzene, with limited production of xylenes as a by-product.

It is another particular object of the present invention to provide a process for the alkylation of reformate with fuel gas or other sources of light olefins to produce an alkylate reformate product containing, inter alia, short chain mono- and dialkylates.

A further object of the present invention is to provide an improved process for
35 the transalkylation of an aromatic compound and a transalkylating agent to produce a

- 5 short chain alkyl aromatic compound, using a transalkylating catalyst comprising a binder-free molecular sieve.

A particular object of the present invention is to provide a process for transalkylating benzene and polyisopropylbenzene or polyethylbenzene to produce cumene or ethylbenzene.

- 10 By way of realizing the foregoing and other objects of the invention, a process for preparing short chain alkyl aromatic compounds is provided which comprises contacting at least one aromatic compound with at least one alkylating agent possessing an aliphatic group having from 1 to 5 carbon atoms under alkylation reaction conditions and in the presence of an alkylation catalyst to provide an alkylated
15 aromatic compound possessing at least one alkyl group derived from said alkylating agent, said alkylation catalyst comprising a binder-free molecular sieve.

- Further provided is a process for preparing short chain alkyl aromatic compounds comprising contacting at least one aromatic compound with at least one transalkylating agent having two or more aliphatic groups each having from 1 to 5
20 carbon atoms under transalkylation reaction conditions and in the presence of a transalkylation catalyst comprising a binder-free molecular sieve.

DESCRIPTION OF SPECIFIC EMBODIMENTS

- 25 The term "aromatic" as used herein includes alkyl substituted and unsubstituted mono- and polynuclear compounds. Compounds of an aromatic character which possess a hetero atom are also useful provided they do not act as catalyst poisons under the reaction conditions selected.

- 30 Substituted aromatic compounds which can be alkylated or transalkylated herein must possess at least one hydrogen atom directly bonded to the aromatic nucleus. The aromatic rings can be substituted with one or more alkyl, aryl, alkaryl, alkoxy, aryloxy, cycloalkyl, halide, and/or other groups which do not interfere with the alkylation or transalkylation reaction.

5 Generally the alkyl groups which can be present as substituents on the aromatic compound contain from about one to 22 carbon atoms and preferably from about one to eight carbon atoms, and most preferably from about one to four carbon atoms.

Suitable alkyl substituted aromatic compounds include toluene, xylenes, isopropylbenzene, normal propylbenzene, alpha-methylnaphthalene, ethylbenzene, cumene, mesitylene, durene, p-cymene, butylbenzene, pseudocumene, o-diethylbenzene, m-diethylbenzene, p-diethylbenzene, isoamylbenzene, isohexylbenzene, pentaethylbenzene, pentamethylbenzene, 1,2,3,4-tetraethylbenzene, 1,2,3,5-tetramethylbenzene, 1,2,4-triethylbenzene, 1,2,3-trimethylbenzene, m-butyltoluene, p-butyltoluene, 3,5-diethyltoluene, o-ethyltoluene, p-ethyltoluene, m-propyltoluene, 4-ethyl-m-xylene, dimethylnaphthalenes, ethylnaphthalene, 2,3-dimethylantracene, 9-ethylantracene, 2-methylantracene, 1-methylantracene, 9,10-dimethylphenanthrene, and 3-methyl-phenanthrene. Higher molecular weight alkylaromatic hydrocarbons can also be used as starting materials and include aromatic hydrocarbons such as are produced by the alkylation of aromatic hydrocarbons with olefin oligomers. Such products are frequently referred to in the art as alkylate and include hexylbenzene, nonylbenzene, dodecylbenzene, pentadecylbenzene, hexyltoluene, nonyltoluene, dodecyltoluene and pentadecyltoluene. Very often alkylate is obtained as a high boiling fraction in which the alkyl group attached to the aromatic nucleus varies in size from about C₆ to C₁₂.

25 Suitable aromatic hydrocarbons include benzene, toluene, xylene, naphthalene, anthracene, naphthacene, perylene, coronene and phenanthrene.

Reformate containing substantial quantities of benzene, toluene and/or xylene constitutes a particularly useful feed for the alkylation process of this invention.

30 The alkylating agents which are useful in the process of this invention generally include any organic compound having at least one alkylating group having from 1 to 5 carbon atoms that is capable of reaction with the aromatic compound. Examples of suitable alkylating agents are olefins such as ethylene, propylene, the butenes and the pentenes; alcohols (inclusive of mono- alcohols, dialcohols, trialcohols, etc.) such as methanol, ethanol, the propanols, the butanols and the pentanols; aldehydes such as

- 5 formaldehyde, acetaldehyde, propionaldehyde, butyraldehyde and n-valeraldehyde; and alkyl halides such as methyl chloride, ethyl chloride, the propyl chlorides, the butyl chlorides and the pentyl chlorides.

Mixtures of light olefins are especially useful as alkylating agents in the alkylation process of this invention. Accordingly, mixtures of ethylene, propylene, butenes and/or pentenes, which are major constituents of a variety of refinery streams, e.g., fuel gas, gas plant off-gas containing ethylene, propylene, etc., naphtha cracker off-gas containing light olefins, refinery FCC propane/propylene streams, etc., are useful alkylating agents herein. For example, a typical FCC light olefin stream possesses the following composition:

15

	Wt. %	Mole %
Ethane	3.3	5.1
Ethylene	0.7	1.2
Propane	14.5	15.3
Propylene	42.5	46.8
Isobutane	12.9	10.3
n-Butane	3.3	2.6
Butenes	22.1	18.32
Pentanes	0.7	0.4

When transalkylation is desired, the transalkylating agent may be a polyalkyl aromatic hydrocarbon containing two or more alkyl groups that each may have from 1 to 5, and preferably 2 to 4 carbon atoms. For example, suitable polyalkyl aromatic hydrocarbons include di-, tri-, and tetraalkyl aromatic hydrocarbons, such as diethylbenzene, triethylbenzene, diethylmethylbenzene (diethyltoluene), diisopropylbenzene, triisopropylbenzene, diisopropyltoluene, dibutylbenzene, and the like. Preferred polyalkyl aromatic hydrocarbons are the dialkyl benzenes. Particularly preferred polyalkyl aromatic hydrocarbons are diisopropylbenzene and diethylbenzene.

25 Reaction products which may be obtained from the process of the invention include ethylbenzene from the reaction of benzene with either ethylene or polyethylbenzenes, cumene from the reaction of benzene with propylene or polyisopropylbenzenes, ethyltoluene from the reaction of toluene with ethylene or

- 5 polyethyltoluenes, cymenes from the reaction of toluene with propylene or polyisopropyltoluenes, and sec-butylbenzene from the reaction of benzene and n-butenes or polybutylbenzenes. The production of cumene from the alkylation of benzene with propylene or the transalkylation of benzene with diisopropylbenzene is another example of the production of a particular product.

- 10 Preferably a synthetic porous crystalline material that is characterized, in its calcined form, by an X-ray diffraction pattern including the following lines, may be used as the binder-free molecular sieve of the present invention:

TABLE A

Interplaner d-Spacing (A)	Relative Intensity, $I/I_0 \times 100$
12.36 \pm 0.4	M-VS
11.03 \pm 0.2	M-S
8.83 \pm 0.14	M-VS
6.18 \pm 0.12	M-VS
6.00 \pm 0.10	W-M
4.06 \pm 0.07	M-S
3.91 \pm 0.07	M-VS
3.42 \pm 0.06	VS

Alternatively, the synthetic porous crystalline material may be characterized by an X-ray diffraction pattern in its calcined form including the following lines:

TABLE B

Interplaner d-Spacing (A)	Relative Intensity, $I/I_0 \times 100$
30.0 \pm 2.2	W-M
22.1 \pm 1.3	W
12.36 \pm 0.4 -	M-VS
11.03 \pm 0.2 -	M-S
8.83 \pm 0.14 -	M-VS
6.18 \pm 0.12 -	M-VS
6.00 \pm 0.10 -	W-M
4.06 \pm 0.07 -	W-S
3.91 \pm 0.07 -	M-VS
3.42 \pm 0.06 -	VS

- 5 More specifically, the calcined form may be characterized by an X-ray diffraction pattern including the following lines:

TABLE C

Interplaner d-Spacing (A)	Relative Intensity, I/I ₀ x 100
12.36 ± 0.4	M-VS
11.03 ± 0.2	M-S
8.83 ± 0.14	M-VS
6.86 ± 0.14	W-M
6.18 ± 0.12	M-VS
6.00 ± 0.10	W-M
5.54 ± 0.10	W-M
4.92 ± 0.09	W
4.64 ± 0.08	W
4.41 ± 0.08	W-M
4.25 ± 0.08	W
4.10 ± 0.07	W-S
4.06 ± 0.07	W-S
3.91 ± 0.07	M-VS
3.75 ± 0.06	W-M
3.56 ± 0.06	W-M
3.42 ± 0.06	VS
3.30 ± 0.05	W-M
3.20 ± 0.05	W-M
3.14 ± 0.05	W-M
3.07 ± 0.05	W
2.99 ± 0.05	W
2.82 ± 0.05	W
2.78 ± 0.05	W
2.68 ± 0.05	W
2.59 ± 0.05	W

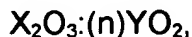
- 5 Most specifically, the synthetic porous crystalline material may be characterized in its calcined form by an X-ray diffraction pattern including the following lines:

TABLE D

Interplaner d-Spacing (A)	Relative Intensity, I/I ₀ x 100
30.0 ± 2.2	W-M
22.1 ± 1.3	W
12.36 ± 0.4	M-VS
11.03 ± 0.2	M-S
8.83 ± 0.14	M-VS
6.86 ± 0.14	W-M
6.18 ± 0.12	M-VS
6.00 ± 0.10	W-M
5.54 ± 0.10	W-M
4.92 ± 0.09	W
4.64 ± 0.08	W
4.41 ± 0.08	W-M
4.25 ± 0.08	W
4.10 ± 0.07	W-S
4.06 ± 0.07	W-S
3.91 ± 0.07	M-VS
3.75 ± 0.06	W-M
3.56 ± 0.06	W-M
3.42 ± 0.06	VS
3.30 ± 0.05	W-M
3.20 ± 0.05	W-M
3.14 ± 0.05	W-M
3.07 ± 0.05	W
2.99 ± 0.05	W
2.82 ± 0.05	W
2.78 ± 0.05	W
2.68 ± 0.05	W
2.59 ± 0.05	W

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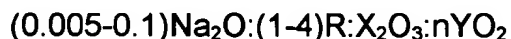
An example of such a synthetic porous crystalline material is MCM-22, which has a composition that satisfies the molar relationship:



wherein X is a trivalent element, such as aluminum, boron, iron and/or gallium,

15 preferably aluminum; Y is a tetravalent element, such as silicon and/or germanium,

5 preferably silicon; and n is at least about 10, usually from about 10 to 150, more usually from about 10 to 60, and even more usually from about 20 to 40. In the as-synthesized form, i.e., before ion exchange, calcination, etc., MCM-22 has a formula, on an anhydrous basis and in term of moles of oxides per n moles of YO_2 , as follows:



10 wherein R is an organic component. The Na and R components are associated with the porous crystalline material as a result of their presence during crystallization, and are easily removed by post-crystallization methods hereinafter more particularly described.

15 MCM-22 is thermally stable and exhibits a high surface area greater than about 400 m²/gm as measured by the BET (Bruenauer, Emmet and Teller) test and unusually large sorption capacity when compared to other crystal structures having similar X-ray diffraction patterns. As is evident from the above formula, MCM-22 is synthesized nearly free of Na cations and thus possesses acid catalyst activity as synthesized. It can, therefore, be used as the alkylation or transalkylation catalyst without having to first undergo an exchange step. To the extent desired, however, the original sodium cations of the as-synthesized material can be replaced in accordance with techniques well known in the art, at least in part, by ion exchange with other cations. Preferred replacement cations include metal ions, hydrogen ions, hydrogen precursor, e.g., ammonium, ions and mixtures thereof. Particularly preferred cations are those which 25 tailor the activity of the catalyst for preparing short chain alkylaromatics. These include hydrogen, rare earth metals and metals of Groups IIA, IIIA, IVA, IB, IIB, IIIB, IVB and VIII of the Periodic Table of the Elements. In its calcined form, MCM-22 appears to be made up of a single crystal phase with little or no detectable impurity crystal phases.

30 MCM-22 can be prepared from a reaction mixture containing sources of alkali or alkaline earth metal cation (M), e.g., sodium or potassium; an oxide of trivalent element X, e.g., aluminum; an oxide of tetravalent element Y, e.g., silicon; an organic (R) directing agent; and water, said reaction mixture having a composition, in terms of mole ratios of oxides, within the following ranges:

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REACTANTS	USEFUL	PREFERRED
YO_2/X_2O_3	10-60	10-40
H_2O/YC_2	5-100	10-50
OH/YO_2	0.01-1.0	0.1-0.5
M/YO_2	0.01-2.0	0.1-1.0
R/YO_2	0.05-1.0	0.1-0.5

In a preferred method of synthesizing MCM-22, the YO_2 reactant contains a substantial amount of solid YO_2 , e.g., at least about 30 wt.% solid YO_2 . Where YO_2 is silica, the use of a silica source containing at least about 30 wt.% solid silica, e.g., Ultrasil (a precipitated, spray dried silica containing about 90 wt.% silica) or HiSil (a precipitated hydrated SiO_2 containing about 87 wt.% silica, about 6 wt.% free H_2O and about 4.5 wt.% bound H_2O of hydration and having a particle size of about 0.02 micron) favors crystal formation from the above mixture. If another source of oxide of silicon, e.g., Q-Brand (a sodium silicate comprised of about 28.8 wt.% of SiO_2 , 8.9 wt.% Na_2O and 62.3 wt.% H_2O) is used, crystallization may yield little if any MCM-22 crystalline material and impurity phases of other crystal structures, e.g., ZSM-12, may be produced. Therefore the YO_2 , e.g., silica, source preferably contains at least about 30 wt.% solid YO_2 , e.g., silica, and more preferably at least about 40 wt.% solid YO_2 .

The organic directing agent for use in synthesizing MCM-22 from the above reaction mixture is hexamethyleneimine.

Crystallization of the MCM-22 crystalline material can be carried out at either static or stirred conditions in a suitable reactor vessel such as, e.g., polypropylene jars or teflon-lined or stainless steel autoclaves. The useful range of temperatures for crystallization is from about 80°C to 225°C for a time sufficient for crystallization to occur at the temperature used, e.g., from about 25 hours to 60 days. Thereafter, the crystals are separated from the liquid and recovered.

Alternatively, MCM-36, MCM-49 or MCM-56 may be used as the synthetic porous crystalline material of the present invention. MCM-36 is described in U.S. Patent Nos. 5,250,277 and 5,292,698, which are incorporated herein by reference. US.

10 5,493,065, 5,371,310 and 5,453,554, which are hereby incorporated by reference.

The synthetic porous crystalline material or molecular sieve herein can also be used in intimate combination with a hydrogenating component such as tungsten, vanadium, molybdenum, rhenium, nickel, cobalt, chromium, manganese, or a noble metal such as platinum, palladium, or platinum/palladium combination where a hydrogenation-dehydrogenation function is to be performed. Such component can be introduced by way of cocrystallization, exchanged into the synthetic porous crystalline material or molecular sieve composition to the extent a Group IIIA element, e.g., aluminum, is in the structure, impregnated therein or physically admixed therewith. Such component can be impregnated in, or on, the composition, for example in the case of platinum, by treating the composition with a solution containing a platinum metal-containing ion. Thus, suitable platinum compounds for this purpose include chloroplatinic acid, platinous chloride and various compounds containing the platinum amine complex. Prior to use, the synthetic porous crystalline material or molecular sieve should be subjected to thermal treatment to remove part or all of any organic constituent present therein.

30 ammonium forms, can be beneficially converted to another form by thermal treatment.

5 treatment, atmospheric pressure is preferred simply for reasons of convenience. The thermal treatment can be performed at a temperature of up to about 925°C.

The stability of the synthetic porous crystalline material of the invention may be increased by steaming. U.S. Patent Nos. 4,663,492; 4,594,146; 4,522,929; and 4,429,176, the entire disclosures of which are incorporated herein by reference,
10 describe conditions for the steam stabilization of the synthetic porous crystalline material herein. The steam stabilization conditions include contacting the catalyst with, e.g., 5-100% steam at a temperature of at least about 300°C (e.g., 300°-650°C) for at least one hour (e.g., 1-200 hours) at a pressure of 101-2500 kPa. In a more particular embodiment, the synthetic porous crystalline material can be made to undergo
15 steaming with 75-100% steam at 315°-500°C and atmospheric pressure for 2-25 hours. In accordance with the steam stabilization treatment described in the above-mentioned patents, the steaming of the synthetic porous crystalline material can take place under conditions sufficient to initially increase the Alpha Value of the catalyst, the significance of which is discussed in U.S. Patent No. 4,992,606, and produce a steamed catalyst
20 having a peak Alpha Value. If desired, steaming can be continued to subsequently reduce the Alpha Value from the peak Alpha Value to an Alpha Value which is substantially the same as the Alpha Value of the unsteamed catalyst.

Prior to its use in the alkylation or transalkylation process of this invention, the synthetic porous crystalline material or molecular sieve should be dehydrated, at least
25 partially. This can be done by heating the synthetic porous crystalline material or molecular sieve to a temperature in the range of from about 200°C to 595°C in an atmosphere such as air, nitrogen, etc. and at atmospheric, subatmospheric or superatmospheric pressures for between about 30 minutes to 48 hours. Dehydration can also be performed at room temperature merely by placing the synthetic porous
30 crystalline or molecular sieve material in a vacuum, but a longer time is required to obtain a sufficient amount of dehydration.

The term "binder-free" as used herein describes the synthetic porous crystalline material or molecular sieve as being substantially free or free of binder materials such as clays or metal oxides, i.e., alumina or silica. The alkylation or transalkylation catalyst

5 according to the present invention may be prepared by mulling crystals of the synthetic porous crystalline material or molecular sieve with water, in the absence of a binder material.

Alkylation of an aromatic compound in accordance with the invention is effected by contact of the reactants at a temperature of between about 0°C and 500°C, and preferably between about 50°C and 250°C. The reaction generally takes place at pressures of from about 0.2 to 250 atmospheres and preferably from about 1 to 70 atmospheres. The molar ratio of aromatic compound to alkylating agent can be from about 0.1:1 to 50:1, and preferably can be from about 0.5:1 to 5:1. Reaction is suitably accomplished utilizing a feed weight hourly space velocity (WHSV) of between about 0.1 hr⁻¹ and 500 hr⁻¹ and preferably from 0.5 hr⁻¹ to 100 hr⁻¹, based upon to total weight of active catalyst.

The alkylation process of this invention can be carried out as a batch-type, semi-continuous or continuous operation utilizing a fixed, fluidized or moving bed catalyst system.

The conditions for transalkylation include a temperature of between about 0°C and 500°C, and preferably between about 160°C and 270°C. The transalkylation pressure ranges from about 0.2 to 250 atmospheres and preferably about 1 to 70 atmospheres. The molar ratio of aromatic compound to transalkylating agent can be from about 0.1:1 to 50:1. Reaction is suitably accomplished using a total feed weight hourly space velocity (WHSV) of between about 0.1 hr⁻¹ and 500 hr⁻¹ and preferably from about 1 hr⁻¹ to 20 hr⁻¹, based upon the total weight of the catalyst.

In order to more fully illustrate the alkylation and transalkylation process of this invention and the manner of practicing the same, the following examples are presented.

30 Comparative Example 1 Preparation of 65 wt.% MCM-22/35 wt.% Al₂O₃ catalyst

The Al₂O₃-bound MCM-22 catalyst was prepared by mulling 65 parts of MCM-22 (the synthesis of MCM-22 is described in U.S. Patent No. 4,954,325 assigned to Mobil Oil Corp, the entire disclosure of which is incorporated herein by reference.) and 35 parts of a pseudoboehmite alumina with water. The mulled powder was formed into

- 5 1/16" extrudates using a screw-type extruder. After drying at 250°F, the extrudate was calcined with nitrogen at 900°F. After cooling to ambient temperature, the extrudate was exchanged with ammonium nitrate to remove residual sodium. After drying, the extrudate was calcined at 1000°F in air. The composition of the MCM-22 after calcination was 65 wt.% MCM-22 and 35 wt.% Al_2O_3 . Catalyst properties are
- 10 summarized in Table 1.

Example 1 Preparation of binder-free MCM-22 catalyst

- The binder-free MCM-22 catalyst according to the present invention was prepared by mulling 100 parts of MCM-22 crystals (the synthesis of which is described
- 15 in U.S. Patent No. 4,954,325) with water. The mulled powder was formed into 1/16" extrudates using a screw-type extruder. After drying at 250°F, the extrudate was exchanged with ammonium nitrate to remove residual sodium. After drying, the extrudate was calcined in nitrogen at 900°F followed by a calcination at 1000°F in air. Catalyst properties are summarized in Table 1.

Table 1. Physical Properties of the Catalysts

22	65 wt. % MCM-22/35 wt. % Al_2O_3	Binder-free MCM-
	Comparative Example 1	Example 1
25	Alpha activity	279
	Real density, g/cc	2.37
	Particle density, g/cc	0.75
30	Surface area, m^2/g	482

Comparative Example 2 Liquid phase ethylbenzene synthesis via benzene-ethylene alkylation over 65 wt.% MCM-22/35 wt.% Al_2O_3 catalyst

- 35 1.0 g of Al_2O_3 -bound MCM-22 catalyst, prepared in Comparative Example 1 as 1/16" extrudates, was chopped to 1/16" length. The catalyst was diluted with sand to 3 cc and charged to an isothermal, down-flow, fixed-bed, 3/8" o.d. reactor. The catalyst was dried at 125°C and 1 atm with 100 cc/min of flowing N_2 for 2 hours. N_2 was turned

5 off. Benzene (reagent grade) was fed into the reactor at 45 WHSV (based on total
catalyst weight) for 1 hr and then at 16.7 WHSV while the reactor temperature and
pressure were increased to 200°C and 500 psig, respectively. After the desired
temperature and pressure were reached, ethylene (Matheson CP grade) was
10 introduced from a mass flow controller at 1.1 WHSV (5.5 benzene/ethylene molar
ratio). After lining out, liquid products were collected in a cold-trap and analyzed off-
line with a Varian 3700 GC. Off-gas was analyzed with an on-line Carle refinery gas
analyzer. Ethylene conversion was determined by measuring unreacted ethylene
relative to feed ethylene. Total material balances were 100±2%. The experiment was
conducted at 200°C, 500 psig, 1.1-2.2 ethylene WHSV, and 5.5 benzene/ethylene
15 molar ratio. Catalyst performance is shown below.

Example 2 Liquid phase ethylbenzene synthesis *via* benzene-ethylene alkylation
over binder-free MCM-22 catalyst

20 1.0 g of binder-free MCM-22 catalyst, prepared in Example 1 as 1/16"
extrudates, was chopped to 1/16" length. The catalyst was tested with the same
procedure described in Comparative Example 2. During the experiment, the effects of
temperature (160-220°C), ethylene WHSV (0.55 to 4.4 h⁻¹) were studied at 500 psig
and 5.5 benzene/ethylene molar ratio. Catalyst performance is shown below.

Comparison of catalyst performance for liquid phase ethylbenzene synthesis

25 Figures 1 and 2 compare catalyst activity for liquid phase ethylbenzene
synthesis at 500 psig, 5.5 benzene/ethylene molar ratio with temperature adjusted
between 160 and 200°C and ethylene WHSV adjusted between 0.55 and 4.4 h⁻¹. To
30 achieve a constant ethylene conversion (e.g., 98%) at 220°C, the binder-free MCM-22
could be operated at 3-4 times higher throughput than the Al₂O₃-bound MCM-22. At
constant throughput, the binder-free MCM-22 provided ≥50°C temperature advantage
over the Al₂O₃-bound MCM-22 to achieve comparable ethylene conversions.

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Figure 1. Ethylbenzene Synthesis with Alumina-Bound MCM-22

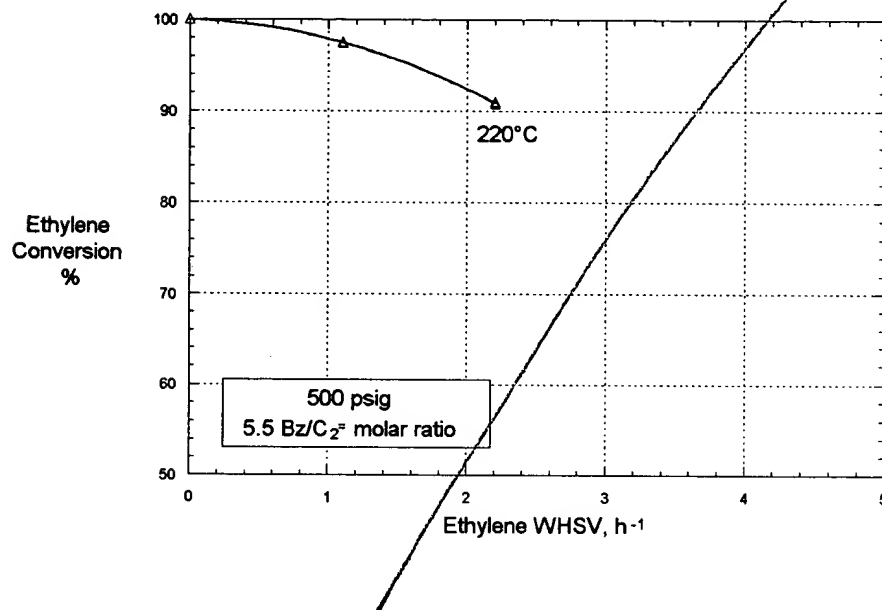
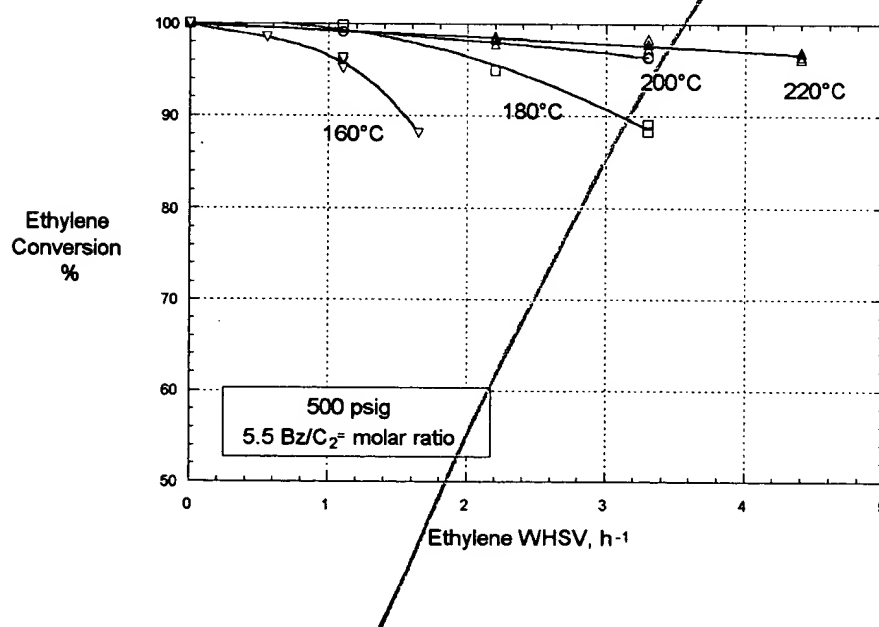


Figure 2. Ethylbenzene Synthesis with Self-Bound MCM-22



- 5 The catalyst performances for liquid phase ethylbenzene synthesis are further compared in Table 2. To achieve comparable conversion at 1.1 ethylene WHSV, the binder-free catalyst could be operated at 161°C vs. 220°C with the Al₂O₃-bound catalyst. To achieve comparable conversion at 220°C, the binder-free catalyst could be operated at 4.4 ethylene WHSV vs. 1.1 ethylene WHSV with the Al₂O₃-bound catalyst.
- 10 catalyst. Ethylbenzene selectivity for the binder-free catalyst, when operated at high throughput, was somewhat lower than the Al₂O₃-bound catalyst. However, the di- and triethylbenzene could be converted to ethylbenzene in the transalkylator by reacting with benzene.

15 **Table 2. Ethylbenzene Synthesis via Benzene/Ethylene Alkylation**

Catalyst	Al ₂ O ₃ -bound MCM-22	Binder-free MCM-22	
Reaction Temperature, °C	220	161	220
Ethylene WHSV	1.1	1.1	4.4
Ethylene Conversion, %	97.4	96.7	96.9
Days on Stream	2.0	5.4	0.4
Product Distribution, wt. %			
Ethylbenzene	93.66	93.36	90.82
Diethylbenzene	6.02	6.22	8.48
Triethylbenzene	0.19	0.24	0.55
Σ	99.87	99.82	99.85
Lights	0.03	0.01	0.00
Xylenes	0.00	0.00	0.00
n-C ₃ -Bz + Cumene	0.00	0.00	0.00
sec-C ₄ -Bz	0.06	0.10	0.08
Heavies	0.04	0.07	0.08
Σ (By Products)	0.13	0.18	0.16

500 psig and 5.5 benzene/ethylene molar ratio.

5

Comparative Example 3 Liquid phase cumene synthesis *via* benzene-propylene alkylation over 65 wt.% MCM-22/35 wt.% Al_2O_3 catalyst

The Al_2O_3 -bound MCM-22 catalyst, prepared in Comparative Example 1 as 1/16" extrudates, was crushed and sized to 30-40 mesh. 0.25 g of this catalyst was diluted with sand to 3 cc and charged to an isothermal, down-flow, fixed-bed, 3/8" o.d. reactor. The catalyst was dried at 125°C and 1 atm with 100 cc/min of flowing N_2 for 2 hours. N_2 was turned off. Benzene was fed into the reactor at 60 cc/hr for 1 hr and then reduced to desired WHSV (based on total catalyst weight) while the reactor temperature and pressure were increased to 120°C and 300 psig, respectively. After reaching 120°C and 300 psig, propylene (Matheson polymer grade) was introduced from a syringe pump at 3 benzene/propylene molar ratio and the temperature was increased to 130°C. After lining out, liquid products were collected in a cold-trap and analyzed off-line with a Varian 3700 GC. Off-gas was analyzed with an on-line Carle refinery gas analyzer. Propylene conversion was determined by measuring unreacted propylene relative to feed propylene. Total material balances were 100±2%. The experiment was conducted at 130°C average reactor temperature, 300 psig, 10-60 propylene WHSV, and 3 benzene/propylene molar ratio. Catalyst performance is shown below.

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Example 3 Liquid phase cumene synthesis *via* benzene-propylene alkylation over binder-free MCM-22 catalyst

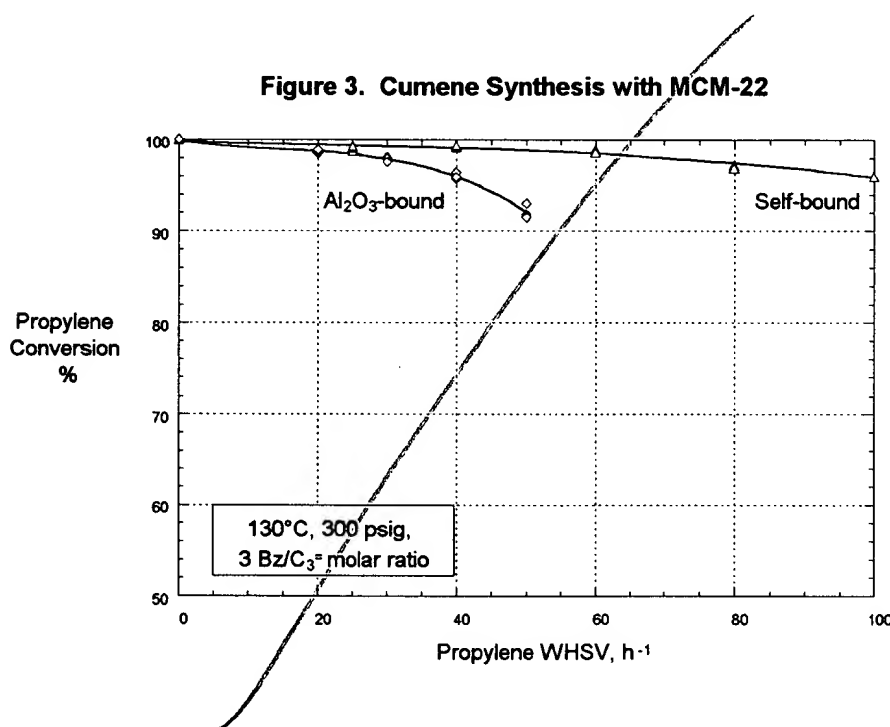
The binder-free MCM-22 catalyst, prepared in Example 2 as 1/16" extrudates, was crushed and sized to 30-40 mesh. 0.1 g of this catalyst was diluted with sand to 3 cc and charged to an isothermal, down-flow, fixed-bed, 3/8" o.d. reactor. The catalyst was tested with the same procedure described in Comparative Example 3. The experiment was conducted at 130°C average reactor temperature, 300 psig, 15-100 propylene WHSV, and 3 benzene/propylene molar ratio. Catalyst performance is shown below.

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Comparison of catalyst performance for liquid phase cumene synthesis

Figure 3 compares catalyst activity for liquid phase cumene synthesis at 130°C average reactor temperature, 300 psig, 3 benzene/propylene molar ratio with propylene WHSV adjusted between 10 and 100 h⁻¹. To achieve a constant propylene conversion (e.g., 96%) at 130°C average reactor temperature, the binder-free MCM-22 could be operated at 100 propylene WHSV when compared to 40 propylene WHSV with the Al₂O₃-bound catalyst.



The catalyst performances for liquid phase cumene synthesis are further compared in Table 3. When operated at 130°C and 40 propylene WHSV, the binder-free MCM-22 offered 3% higher propylene conversion than the Al₂O₃-bound MCM-22. To achieve 96% propylene conversion at 130°C, the binder-free MCM-22 can be operated at 100 propylene WHSV vs. 40 propylene WHSV with the Al₂O₃-bound catalyst. Although cumene selectivity was somewhat lower with the binder-free MCM-22, the overall alkylation selectivity was higher due to its lower propylene oligomers

- 5 make. Propylene oligomers are rejected as by-products which reduce net propylene conversion to cumene. The di- and triisopropylbenzene can be converted to cumene in the transalkylator by reacting with benzene.

10 **Table 3. Cumene Synthesis via Benzene/Propylene Alkylation**

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Catalyst	Al ₂ O ₃ -bound MCM-22 catalyst	Binder-free MCM-22	
Propylene WHSV	40.0	40.0	100.0
15 Propylene Conversion, %	95.9	99.2	96.0
Days on Stream	3.3	3.5	6.2
Product Distribution, wt. %			
Cumene	87.03	85.73	85.65
Diisopropylbenzene	10.94	12.39	12.37
20 <u>Triisopropylbenzene+</u>	<u>1.01</u>	<u>1.17</u>	<u>1.11</u>
Σ	98.98	99.29	99.13
Propylene Oligomers	1.01	0.69	0.84
25 <u>n-C₃-Bz</u>	<u>0.02</u>	<u>0.02</u>	<u>0.02</u>
Σ (By Products)	1.03	0.71	0.86

130°C average reactor temperature, 300 psig, & 3 benzene/propylene molar ratio.

30 Comparative Example 4. Liquid phase ethylbenzene synthesis via benzene/
polyethylbenzene transalkylation over 65 wt. % MCM-
22/35% wt. % Al₂O₃ catalyst

- 35 0.5 g of Al₂O₃-bound MCM-22 catalyst, prepared in Comparative Example 1 as 1/16" extrudates, was chopped to 1/16" length. The catalyst was diluted with sand to 3 cc and charged to an isothermal, down-flow, fixed-bed, 3/8" o.d. reactor. The catalyst was dried at 125°C and 1 atm with 100 cc/min flowing N₂ for 2 hours. N₂ was turned off and reactor pressure was set to 500 psig by a grove loader. The feed (74.94 wt. %
- 40 benzene, 23.04 wt. % diethylbenzenes, 0.14 wt. % butylbenzenes, and 1.60 wt. % triethylbenzenes) was fed into the reactor at 60 cc/hr for 1 hour and then at 5 total WHSV. The reactor temperature was then increased to 260°C. After lining out, liquid

5 products were collected in a cold-trap and analyzed off-line with an HP 5890 GC. The catalyst was then tested at 10 WHSV. Total material balances were $100 \pm 2\%$. Catalyst performance is shown below.

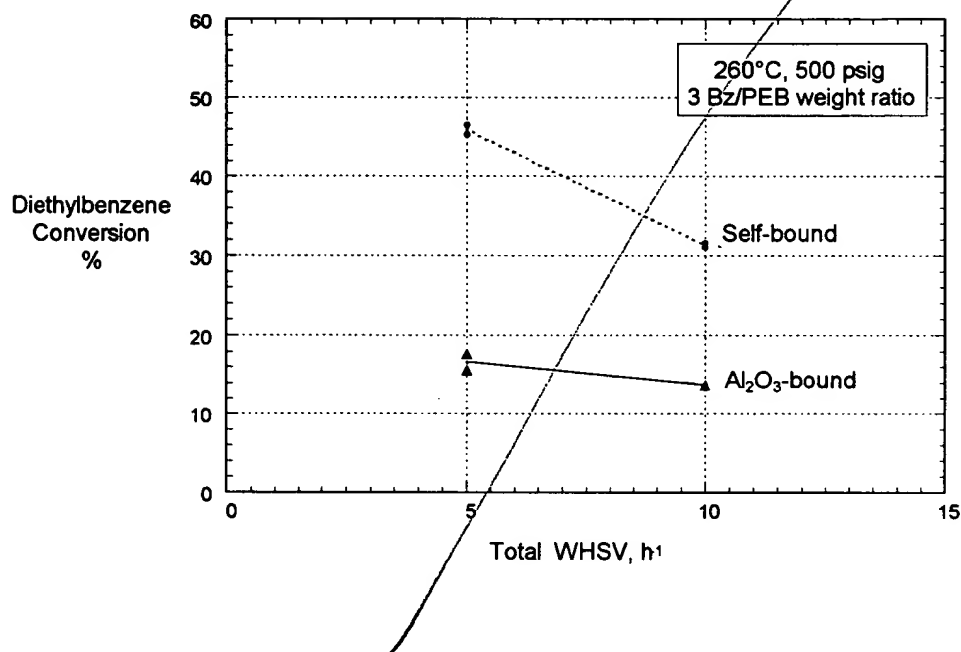
10 Example 4. Liquid phase ethylbenzene synthesis *via* benzene/polyethylbenzene transalkylation over binder-free MCM-22 catalyst

0.5 g of binder-free MCM-22 catalyst, prepared in Example 1 as 1/16" extrudates, was chopped to 1/16" length. The catalyst was tested with the same
15 procedure described in Comparative Example 4. Catalyst performance is shown below.

Comparison of catalyst performance for liquid phase ethylbenzene synthesis *via* benzene/polyethylbenzene transalkylation

20 Figure 4 compares catalyst activity for liquid phase ethylbenzene synthesis *via* benzene/polyethylbenzene transalkylation at 260°C, 500 psig, 3:1 benzene/polyethylbenzene weight ratio with total WHSV adjusted between 5 and 10 hr⁻¹. The binder-free MCM-22 was significantly more active than the Al₂O₃-bound MCM-22.
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Figure 4. Ethylbenzene Synthesis via Benzene/PEB Transalkylation with MCM-22



The catalyst performances of benzene/polyethylbenzene transalkylation are further compared in Table 4. When operated at 5 total WHSV, the binder-free MCM-22 offered about 3-fold higher diethylbenzene, butylbenzene, and triethylbenzene conversions than the Al₂O₃-bound MCM-22. At these higher conversions, the binder-free MCM-22 provided 3% higher ethylbenzene selectivity than the Al₂O₃-bound MCM-22.

5 **Table 4. Ethylbenzene synthesis via benzene/polyethylbenzene transalkylation**

Catalyst		Al ₂ O ₃ -bound MCM-22	Binder-free MCM-22
10	Diethylbenzene Conv. %	15.6	45.3
	Butylbenzene Conv. %	7.9	20.4
	Triethylbenzene Conv. %	14.3	42.0
	Hours on Stream	47.5	29.3
<hr/>			
Product Selectivity, %			
15	Lights	0.40	0.14
	Toluene	0.13	0.07
	Ethylbenzene	95.17	98.39
	Xylenes	0.00	0.00
	Styrene	0.00	0.00
20	C9 Aromatics	0.30	0.15
	Others	4.00	1.25

25 **260°C, 500 psig, 5 total WHSV, and 3:1 benzene/polyethylbenzene weight ratio (74.94 wt.% benzene, 23.04 wt.% diethylbenzenes, 0.14 wt. % butylbenzenes, and 1.60 wt. % triethylbenzenes).**

30 Comparative Example 5 Liquid phase cumene synthesis via benzene/diisopropylbenzene transalkylation over 65 wt. % MCM-22/35% wt. % Al₂O₃ catalyst

35 0.5 g of Al₂O₃-bound MCM-22 catalyst, prepared in Comparative Example 1 as 1/16" extrudates, was chopped to 1/16" length. The catalyst was diluted with sand to 3 cc and charged to an isothermal, down-flow, fixed-bed, 3/8" o.d. reactor. The catalyst was dried at 125°C and 1 atm with 100 cc/min flowing N₂ for 2 hours. N₂ was turned off and reactor pressure was set to 300 psig by a grove loader. The feed (75.0 wt. % benzene, 8.3 wt. % m-diisopropylbenzene and 16.7 wt. % p-diisopropylbenzene) was fed into the reactor at 60 cc/hr for 1 hour and then at 4 total WHSV (based on total catalyst weight). The reactor temperature was then increased to 200°C. After lining out, liquid products were collected in a cold-trap and analyzed off-line with an HP 5890

- 5 GC. The catalyst was further tested at 210°C and 220°C under otherwise identical conditions. Total material balances were 100±2%.

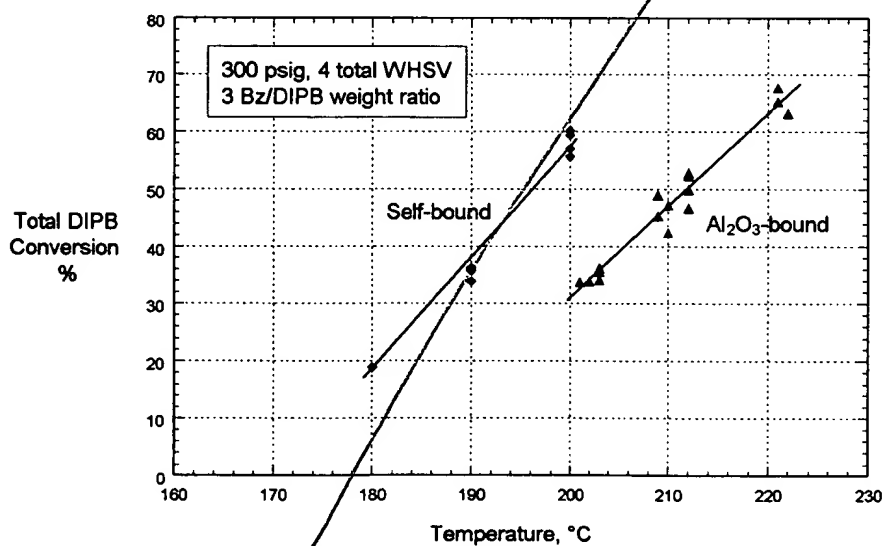
Example 5 Liquid phase cumene synthesis via benzene/diisopropylbenzene
transalkylation over binder-free MCM-22 catalyst

10 1.0 g of binder-free MCM-22 catalyst, prepared in Example 1 as 1/16" extrudates, was chopped to 1/16" length. The catalyst was tested at 180°C, 190°C, and 200°C with the same procedure described in Comparative Example 5. Catalyst performance is compared with that of Al₂O₃-bound MCM-22.

15 Comparison of catalyst performance for liquid phase cumene synthesis via benzene/diisopropylbenzene transalkylation

20 Figure 5 compares catalyst activity for liquid phase cumene synthesis via benzene/diisopropylbenzene transalkylation at 300 psig, 4 total WHSV, and 3:1 benzene/diisopropylbenzene weight ratio.

Figure 5. Cumene Synthesis via Benzene/DIPB Transalkylation with MCM-22



The catalyst performances for benzene/diisopropylbenzene transalkylation are tabulated in Table 5. At ~36% diisopropylbenzene conversion, binder-free MCM-22
25 can be operated at 190°C vs. 203°C with the Al₂O₃-bound MCM-22. At ~200°C, binder-

- 5 free MCM-22 had 67% higher diisopropylbenzene conversion (59.3% vs. 35.6%) when compared with Al₂O₃-bound MCM-22.

Table 5. Cumene synthesis via benzene/diisopropylbenzene transalkylation

10	Catalyst	Al ₂ O ₃ -bound MCM-22	Binder-free MCM-22	
	Temperature, °C	203	190	200
	Diisopropylbenzene Conv. %	35.6	35.8	59.3
	Days on Stream	3.9	1.0	3.0
15	Product Selectivity, wt%			
	Lights (C ₇ -)	0.00	0.02	0.03
	Ethylbenzene	0.11	0.15	0.15
	Cumene	98.84	98.67	98.97
20	n-C ₃ -Bz	0.07	0.06	0.10
	Triisopropylbenzene	0.51	0.69	0.51
	Others	0.47	0.41	0.24

300 psig, 4 total WHSV (based on total catalyst weight), and 3:1 benzene/diisopropylbenzene weight ratio.